A novel approach for stabilizing fresh urine by calcium hydroxide addition

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1. Introduction

Source separation of human excreta is a resource-efficient alternative to conventional water-borne urban drainage and wastewater treatment (Larsen et al., 2009; Larsen and Gujer, 2001). In fast growing cities, especially in mid- and low-income countries, there is a large potential for source-separating systems (Gounden et al., 2006; Huang et al., 2007; Medilanski et al., 2006), but its implementation is also favorable in developed countries (Larsen et al., 2013). One important aspect of source separation is the recovery of nutrients from urine: it allows the recycling of nutrients to agriculture, prevents environmental pollution and gives the opportunity to recover financial value by selling the nutrients as fertilizer (Udert et al., 2015). To maximize the recovery of nutrients from urine and to prevent malodor, urine has to be stabilized. Stabilization mainly means preventing enzymatic urea hydrolysis.

In this study, we investigated the prevention of enzymatic urea hydrolysis in fresh urine by increasing the pH with calcium hydroxide (Ca(OH)₂) powder. The amount of Ca(OH)₂ dissolving in fresh urine depends significantly on the composition of the urine. The different urine compositions used in our simulations showed that between 4.3 and 5.8 g Ca(OH)₂ dissolved in 1 L of urine at 25 °C. At this temperature, the pH at saturation is 12.5 and is far above the pH of 11, which we identified as the upper limit for enzymatic urea hydrolysis. However, temperature has a strong effect on the saturation pH, with higher values being achieved at lower temperatures. Based on our results, we recommend a dosage of 10 g Ca(OH)₂ L⁻¹ of fresh urine to ensure solid Ca(OH)₂ always remains in the urine reactor which ensures sufficiently high pH values. Besides providing sufficient Ca(OH)₂, the temperature has to be kept in a certain range to prevent chemical urea hydrolysis. At temperatures below 14 °C, the saturation pH is higher than 13, which favors chemical urea hydrolysis. We chose a precautionary upper temperature of 40 °C because the rate of chemical urea hydrolysis increases at higher temperatures but this should be confirmed with kinetic studies. By considering the boundaries for pH and temperature developed in this study, urine can be stabilized effectively with Ca(OH)₂ thereby simplifying later treatment processes or making direct use easier.

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effective, they all have flaws, especially for on-site application: the addition of acid is potentially dangerous, requires pumping equipment and exact dosing; electrochemical treatment only allows for short-term inhibition of enzymatic urea hydrolysis and does not prevent later contact with urease-active microorganisms; finally, partial nitrification is a complex biological process that requires process regulation—at least at the present level of development (Fumassoli et al., 2016).

An alternative and potentially more elegant stabilization method would be the inhibition of enzymatic urea hydrolysis at high pH values. Urease activity is negatively affected not only by a low pH but also by a high pH. The pH optima for bacterial urease have been reported to be in a range of 6.8–8.7 (Mobley and Hausinger, 1989). In contrast to strong acids, several strong bases are available as basic salts such as calcium and magnesium oxides and hydroxides. These basic salts are advantageous for several reasons: firstly, they are commonly used to enhance soil quality—primarily in order to increase the pH of acidic soils, but also to provide essential plant nutrients such as calcium and magnesium (Shaul, 2002). Secondly, calcium and magnesium form phosphate minerals at high pH values, allowing the separation of phosphates. Finally, there is a good possibility that the high pH of the solution will reduce the number of pathogens considerably, for example viruses and helminthes (Eriksen et al., 1996; Pancorbo et al., 1988).

However, urea is not only degraded by the enzyme urease. At high pH values and high temperatures, urea also decomposes chemically (Warner, 1942) into ammonia and carbon dioxide. This process has to be considered when stabilizing urine at high pH values.

In this study we assessed the feasibility of preventing enzymatic urea hydrolysis in fresh urine by three common calcium compounds: calcium oxide (CaO), calcium hydroxide (Ca(OH)₂) and calcium carbonate (CaCO₃). We chose these compounds because they are cheap (US$ 0.08 kg⁻¹ for Ca(OH)₂) (Muster et al., 2013) and because CaO and Ca(OH)₂ are commonly used in sanitation for disinfection purposes (Ronteltap et al., 2014). Another benefit of Ca(OH)₂ is that calcium phosphate precipitates during the stabilization process which can be recovered and used for fertilizer production (Rittmann et al., 2011). We also developed a preliminary design chart showing conditions where urea loss will likely be minimized.

2. Materials and methods

2.1. Simulation of chemical speciation

The software PHREEQC version 3.1.7.9213 (Parkhurst and Appelo, 1999) was used to simulate the speciation of chemical compounds at thermodynamic equilibrium. The components used for the chemical speciation calculations are shown in Table 1. Simulations were conducted for different urine compositions (Table 2). For ionic strength corrections, we used the Pitzer approach implemented in a database described in (Wächter et al. in prep.). The composition of fresh urine was based on results from real human urine. Urine samples U1, U2, U3 and U4 were characterized in this study while the composition of urine samples U5 (Udert et al., 2003b) and U6 (Udert et al., 2006) were obtained from literature. All the compounds in the lower part of Table 2 were used in the simulations. Acetate was used as a proxy for chemical oxygen demand (COD), under the assumption that 1 g COD is equivalent to 1.08 g of acetate (Grady et al., 1999).

2.2. Urine solutions

For all experiments described in this paper, fresh male and female human urine was collected in glass bottles (10 persons, age 25–50). For longer experiments, the urine samples were stored at 4 °C before use. Only urine with a pH lower than 6.9 was used, because higher pH values were an indication of urea hydrolysis and therefore contamination with urease-active microorganisms (Udert et al., 2003b). Urine samples U1a and U4 (Table 2) depict average compositions of the fresh urine used in the experiments for this study.

2.3. Dissolution of Ca(OH)₂ in fresh urine

A dosage of 5 g Ca(OH)₂ (102047 Calcium hydroxide, Merck, Darmstadt, Germany) was added to 1 L of fresh urine to investigate the dissolution. The dosage was based on results from preliminary PHREEQC simulations, which showed that 4.3 g Ca(OH)₂ is the maximum solubility for the specific urine composition (U1) used in the experiments at 25 °C. During the dissolution of Ca(OH)₂ in fresh urine, pH was measured with 5 s intervals (Sentix 41, WTW, Weilheim, Germany). The initial pH of the urine was 6.63.

2.4. Ideal pH range for inhibition of enzymatic urea hydrolysis

For the experiments to determine the ideal pH range for the inhibition of enzymatic urea hydrolysis, nine 250 mL glass bottles (Schott, borosilicate glass 3.3, Mainz, Germany) were filled with 250 mL unfiltered urine (U4). The bottles were washed thoroughly, with hot water and detergents to minimize the presence of urease. The bottles were also stored closed with lids. To all bottles, 37 µL urease (116493 Urease, Merck, Darmstadt, Germany) was added to ensure a urea hydrolysis rate of 3000 g N m⁻³ d⁻¹ which is typical for urine storage tanks (Udert et al., 2003b). The pH of the eight bottles was adjusted with 10 mol L⁻¹ NaOH (Merck, Darmstadt, Germany) at different pH levels between 9.0 and 12.5 in increments of 0.5 pH units. In one bottle the pH was not adjusted and served as a blank. All bottles were placed in the dark and agitated at 150 rpm using a plate shaker (SM-30, Edmund Bühler GmbH, Hechingen, Germany). The duration of the experiment was 10 days and the ammonium concentration was measured and the pH recorded. For analysis, each bottle was placed on a magnetic stirrer at 500 rpm and a 1 mL unfiltered sample was taken once the pH was stable. The pH of each bottle was controlled daily and readjusted with 10 mol L⁻¹ NaOH, if required. The pH probe was calibrated weekly with pH 4 and pH 7 buffer solutions.

2.5. Long-term stabilization of fresh urine with Ca(OH)₂

The long-term behavior of urine dosed with Ca(OH)₂ was investigated for the addition of varying amounts of Ca(OH)₂ to 1 L of fresh urine (U1a): 2.5 g, 5 g and 10 g. Urease (116493 Urease, Merck, Darmstadt, Germany) was added to the batches in order to obtain a urea hydrolysis rate of 3000 g N m⁻³ d⁻¹. Two control experiments were also conducted, one without adding either urease or Ca(OH)₂ and one with urease but no Ca(OH)₂. All batches were covered in order to reduce ammonia volatilization and continuously stirred. Total ammonia concentrations were regularly measured (cuvette test for ammonium, Hach-Lange, Düsseldorf, Germany) and pH (340, WTW, Weilheim, Germany) was measured whenever a sample was taken. The stabilization experiment was conducted for 27 days.

2.6. Stabilization of consecutive batches of fresh urine with Ca(OH)₂

An initial deposit of 5 g Ca(OH)₂ was provided followed by additional batches of fresh urine (U1a) (250 ml) after 24, 48 and 72 h. The mixture was continuously stirred magnetically. All portions of fresh urine, which were added to the reactor, had a pH...
value below 6.9. During the experiment, the batch was covered to avoid volatilization of ammonia and the pH was measured shortly before and after adding a new portion of fresh urine.

2.7. Analytical methods

Concentrations of dissolved compounds were determined with ion chromatography (IC 881 Compact IC pro, Metrohm, Zofingen, Switzerland) and inductively coupled plasma-optical emission spectrometry (ICP-OES, Ciros, Spectro Analytical Instruments, Kleve, Germany). The standard deviation for all wet chemical measurements was 5% or less.

2.8. Analysis of precipitates

Urine and Ca(OH)$_2$ were mixed with a magnetic stirrer, left overnight without stirring and centrifuged to separate solids from the liquid fraction (10 min at 4000 rpm). The separated solids were dried under a fume hood for several days. X-ray diffractometry (XRD) (Power X-ray diffractometer AXS D8 Advance, Bruker, Coventry, UK) was used to determine the composition of the precipitates.

### Table 1

<table>
<thead>
<tr>
<th>Formula</th>
<th>Name</th>
<th>Formula</th>
<th>Name</th>
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</thead>
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<td>CaCl$_2$ 6H$_2$O</td>
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<tr>
<td>(NH$_4$)$_2$SO$_4$</td>
<td>–</td>
<td>K$_2$SO$_4$</td>
<td>arcanite Mg(OH)$_2$ brucite</td>
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<tr>
<td>NH$_4$MgPO$_4$·6H$_2$O</td>
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<td>K$_2$HPO$_4$</td>
<td>mercalite Mg(OH)Cl –</td>
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<td>CaO</td>
<td>line</td>
<td>NaCl</td>
<td>MgCl$_2$ chloromagnesite</td>
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<td>portlandite</td>
<td>NaOH</td>
<td>MgCl$_2$ 2H$_2$O –</td>
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<tr>
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<td>hydrophilite</td>
<td>Na$_2$SO$_4$</td>
<td>misenite MgCl$_2$ 4H$_2$O –</td>
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<td>mirabilite MgCl$_2$·2H$_2$O MgCl$_2$ 6H$_2$O –</td>
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<td>K$_2$Mg(SO$_4$)$_2$ langbeinite MgSO$_4$·4H$_2$O leonardite</td>
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<td>KCl</td>
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### Table 2

Composition of different fresh urine samples. U1a is the fresh urine before stabilization with Ca(OH)$_2$ while U1b is the composition of the same urine after stabilization. The compositions U1, U2, U3 and U4 were measured in this study while U5 (Udert et al., 2006) and U6 (Udert et al., 2003b) were taken from literature. The composition of US is based on medical values. The columns marked with (sim.) for U1a and U1b show the values used for PHREEQC simulations. All other compositions were added as is to PHREEQC.

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<th>Measurement</th>
<th>Unit</th>
<th>U1a (meas.)</th>
<th>(sim.)</th>
<th>U1b (meas.)</th>
<th>(sim.)</th>
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<th>U3</th>
<th>U4</th>
<th>U5</th>
<th>U6</th>
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<td></td>
<td>33</td>
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<td>–</td>
<td>–</td>
<td>–</td>
<td>–</td>
<td>–</td>
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<tr>
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<td>5460</td>
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<td>–</td>
<td>–</td>
<td>–</td>
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<tr>
<td>NO$_2$-N</td>
<td>g m$^{-3}$</td>
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<td></td>
<td>&lt;10</td>
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<td>–</td>
<td>–</td>
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<td>g m$^{-3}$</td>
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<td></td>
<td>10</td>
<td></td>
<td>–</td>
<td>–</td>
<td>–</td>
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<td>–</td>
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<td>PO$_4$-P</td>
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<td>260</td>
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<td>436</td>
<td>419</td>
<td>437</td>
<td>346</td>
<td>418</td>
<td>170</td>
<td>480</td>
<td>386</td>
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<td>Dissolved COD</td>
<td>g m$^{-3}$</td>
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### 3. Results

3.1. Simulation of Ca(OH)$_2$ solubility in fresh urine

In Fig. 1A–C the simulated percentage dissolution and the resulting pH at 25 °C are shown for different dosage concentrations of CaO, Ca(OH)$_2$ and CaCO$_3$ added to fresh urine. Calcium oxide is highly soluble, but immediately transforms to Ca(OH)$_2$, giving rise to identical results as for Ca(OH)$_2$. Both Ca(OH)$_2$ and CaCO$_3$ are sparingly soluble: a saturated solution is rapidly achieved for all of the three minerals: at 3.5 g CaO L$^{-1}$, 4.3 g Ca(OH)$_2$ L$^{-1}$ and 2.0 g CaCO$_3$ L$^{-1}$. However, the addition of CaCO$_3$ only leads to a pH of 7.2, whereas the addition of CaO and Ca(OH)$_2$ results in a pH of 12.5. As typical pH optima for urease are between 6.8 and 8.7 (Mobley and Hausinger, 1989), CaCO$_3$ dosage is unsuitable for the inhibition of enzymatic urea hydrolysis.

Since CaCO$_3$ is not suitable for urease inhibition and CaO is immediately transformed into Ca(OH)$_2$, the theoretical saturation pH as a function of temperature was only assessed for the latter (Fig. 1D). The saturation pH decreases as the temperature increases. The urine curve represents the saturation pH values for all the compositions given in Table 2. The saturation pH values were the same even though the compositions differed substantially. In fact, even in pure water, the saturation pH values of Ca(OH)$_2$ is similar to
that of urine (Fig. 1D).

Fig. 2 shows the amount of Ca(OH)$_2$ that dissolves in different urine compositions as a function of temperature and varies significantly depending on the urine composition and, to a lesser extent, the temperature. For example, the composition used in the experimental work of this study (U1) resulted in a simulated value of 4.3 g Ca(OH)$_2$ L$^{-1}$ dissolving at 25 $^\circ$C while 5.80 g Ca(OH)$_2$ L$^{-1}$ dissolved for composition U5 at 25 $^\circ$C. The simulation results (Fig. 2) also illustrate that Ca(OH)$_2$ has an inverse solubility which means that less solid dissolves at higher temperatures.

3.2. Dissolution of Ca(OH)$_2$ in fresh urine

The experimental addition of 5 g Ca(OH)$_2$ L$^{-1}$ of urine at 25 $^\circ$C led to comparable results for the simulated and measured composition of stabilized urine (Table 2): dissolved phosphate and magnesium were nearly completely removed from the liquid phase, whereas the concentration of dissolved calcium increased. All other parameters, with the exception of pH, are only marginally affected by the addition of Ca(OH)$_2$.

In the experiments with 5 g Ca(OH)$_2$, a stable high pH value was reached within less than 1 min (Fig. 3). A Ca(OH)$_2$ deposit could

![Fig. 1. Simulated pH (—) and percentage of dissolution (—) for the addition of CaO (A), Ca(OH)$_2$ (B) and CaCO$_3$ (C) to fresh urine at 25 $^\circ$C and saturation pH of Ca(OH)$_2$ in urine (—) and water (—) as a function of temperature (D). When CaO is dissolved in water, it converts to Ca(OH)$_2$. For this reason, the dissolution of CaO as well as the dissolution of Ca(OH)$_2$ are depicted in A.](image1)

![Fig. 2. Simulated amount of Ca(OH)$_2$ dissolving in different urine compositions: U1 (—), U2 (—), U3 (—), U4 (——), U5 (——) and U6 (——). The compositions are given in Table 2.](image2)

![Fig. 3. Experimentally observed pH increase after the addition of 5 g Ca(OH)$_2$ to 1 L of fresh urine.](image3)
keep the pH of consecutively added fresh urine between 12 and 12.5 during the experimental duration of 6 days (Fig. 4). For all laboratory experiments, the measured pH increase caused by the addition of Ca(OH)\(_2\) was in line with the results of the computer simulations: a pH around 12.5 was found for dosages of at least 5 g Ca(OH)\(_2\) per liter.

### 3.3. Experimentally determined urine stability at high pH values

Calcium hydroxide addition can prevent urea hydrolysis for a long period, in our experiments, 27 days (Fig. 5). In the case of pure fresh urine without the addition of urease or Ca(OH)\(_2\), there is only a slow increase in ammonium leading to a doubling of the ammonium concentration after 27 days, which implies that the urine contained few urease-active microorganisms (Fig. 5A). When the untreated urine was spiked with urease, the ammonium concentration increased rapidly: within the first 10 days of the experiment, the concentration increased to about 4300 mg N L\(^{-1}\) (Fig. 5B), which is more than tenfold of the initial concentration (340 mg N L\(^{-1}\)). As depicted in Fig. 5C, the ammonium concentration of urine spiked with Ca(OH)\(_2\) remained within the initial range. Fig. 5D shows that the pH in fresh urine spiked with urease reaches a value of 9.3 within the first day. In contrast, the pH of pure fresh urine slowly increased to 7.7 within the 27 days of the experiment. For urine spiked with Ca(OH)\(_2\), a pH value of 12.4 was reached with 5 g Ca(OH)\(_2\) per liter or more. However, with an addition of 2.5 g Ca(OH)\(_2\) per liter, the pH value stayed at 11.9, which indicates that the solution was not saturated with Ca(OH)\(_2\).

A set of experiments using urine composition U4 and fixed pH values was conducted in order to determine, how far the pH value has to be increased to inhibit enzymatic urea hydrolysis. The production of total ammonia during ten days was used as a measure of the extent of enzymatic urea hydrolysis (Fig. 6). The total ammonia concentration did not change substantially, when the pH value was 11 or higher. At lower pH values, the production of ammonia increased with decreasing pH values. The pH adjustments were done with NaOH and not calcium compounds for these experiments.

### 3.4. Calculation of non-enzymatic urea hydrolysis

Other researchers have shown that chemical urea hydrolysis...
strongly depends on pH (Warner, 1942) and temperature (Chin and Kroontje, 1963; Warner, 1942) as shown in Fig. 7. According to the literature data, the half-life of urea is in the range of days below 66 °C, whereas at 100 °C, the half-life is reduced to hours (Fig. 7A and B). The pH value also influences urea hydrolysis, but only at the extreme ends that is below pH 2 and above pH 12 (Fig. 7A). At pH values higher than 12, the half-life is decreasing, though a significant reduction of the half-life is only observed, when the pH values rise above 13. Fig. 7B shows the impact of temperature on urea, where an increased chemical hydrolysis is favored by increasing temperature.

### 3.5. Simulation of phosphate precipitation

Our experimental results showed that the addition of Ca(OH)$_2$ caused complete precipitation of phosphate and magnesium as indicated by the decrease in concentration (Table 2). It is known that calcium phosphate compounds such as apatite precipitate at high pH values (Nriagu et al., 1984). However, XRD analysis of the precipitates revealed only the presence of the calcium compounds Ca(OH)$_2$ and CaCO$_3$, and not of any crystalline calcium phosphate minerals. Also, no struvite pattern was found, although struvite typically precipitates from urine at elevated pH values ( Udert et al., 2003a). Computer simulations of the chemical speciation indicate that hydroxyapatite (Ca$_5$(PO$_4$)$_3$(OH)) should be present at thermodynamic equilibrium (Fig. 8). However, amorphous calcium phosphate (usually not detected by XRD analysis) is often the first compound to precipitate ( Nancollas and Zhang, 1994) and it takes some time to transform to Ca$_5$(PO$_4$)$_3$(OH). According to the computer simulations, the magnesium found in urine at these conditions precipitates as Mg(OH)$_2$. For dosages of Ca(OH)$_2$ around 4.3 g and higher, an increase in Ca(OH)$_2$ can be observed due to incomplete dissolution of the compound.

### 4. Discussion

#### 4.1. Urine stabilization with Ca(OH)$_2$

We have developed a tentative design chart (Fig. 9) that can be used to determine the ideal operating conditions with respect to pH and temperature for the stabilization of urine with Ca(OH)$_2$. The solubility of Ca(OH)$_2$ defines the upper boundary of the design chart. The saturation pH does not depend on the composition of urine (Fig. 1D) but the amount of Ca(OH)$_2$ required to create a saturated solution does depend on the composition of the urine (Fig. 2).

For example, the experimental results showed that a dosage of 5 g Ca(OH)$_2$ was enough to keep urine composition U1 stabilized since the saturation concentration was 4.3 g Ca(OH)$_2$ L$^{-1}$ of urine. However, a dosage of 5 g L$^{-1}$ of urine would not saturate urine compositions U5 and U6. The complex aqueous chemistry and variability in compositions between urine samples results in different Ca(OH)$_2$ solubilities in fresh urine. For this reason, we recommend a dosage of 10 g Ca(OH)$_2$ L$^{-1}$ of fresh urine because this will effectively stabilize a wide range of urine solutions with different compositions while still retaining residual solid Ca(OH)$_2$ in solution to ensure saturation (Fig. 4).

The design chart (Fig. 9) also shows that for temperatures below 14 °C, a saturation pH of 13 (point (2)) or higher is obtained. At these low temperatures the rate of chemical urea hydrolysis will be slow but the excessively high pH values will increase the rate significantly. As a first guess, we suggest a minimum temperature of 14 °C as a limit. A more accurate limit will have to be determined with specific kinetic experiments.

The shaded green region in the design chart is a region where urea loss is low because both enzymatic urea hydrolysis and chemical urea hydrolysis are limited. A saturated Ca(OH)$_2$ solution would operate along line 2–3 which would be the ideal conditions of operation. Technically, operation could also occur below this line, in the green area, but operation in this region would be challenging because it would require active dosing of Ca(OH)$_2$ or another alkaline chemical using dosing pumps or solid dosing devices and additional equipment. Our approach of passive dosing with a fixed deposit of Ca(OH)$_2$ powder, in excess of that required for saturation, is much easier because no active pH control is then required. We also have to ensure that Ca(OH)$_2$ is added to fresh urine as soon as possible, ideally in a stabilisation tank below the toilet or urinal so that the piping between the user interface and the stabilization tank is short and the retention time is too low for substantial urea hydrolysis to occur.

The transparent rectangular region demarks conditions which favor enzymatic urea hydrolysis. Based on our finding, the pH value in fresh urine has to be 11 or higher to suppress urease activity at a temperature of 25 °C. At higher temperatures and pH values below 11 there is a transition from enzymatic to chemical urea hydrolysis. Operating above a temperature of 55 °C inactivates the enzyme urease (Lencki et al., 1992) therefore chemical urea hydrolysis is the only reason for urea loss above this temperature.

Excessively high temperatures have to be avoided because this promotes the chemical hydrolysis of urea. We have introduced a tentative and conservative border for safe urea stabilization at 40 °C. The exact upper temperature limit has to be determined as a function of hydrolysis kinetics, duration of urine storage, acceptable ammonia losses and the desirable safety factor. Further research into the kinetic rates of chemical urea hydrolysis in urine is therefore needed as this will help to better define this temperature boundary.

#### 4.2. Alternative calcium salts for dosing

Calcium carbonate dissolution will cause a pH increase in fresh urine, but the resulting pH is too low (less than 7.2) to inhibit urease activity. Therefore, CaCO$_3$ is not suitable for urine stabilization. Calcium oxide is also not suited for urine stabilization but for other reasons. The addition of CaO to fresh urine leads to pH levels...
around 12, which is sufficient for suppressing the enzymatic conversion of urea to ammonia under ambient temperatures. However, when CaO is dissolved in aqueous solutions, such as urine, a vigorous release of heat of 1.14 MJ kg\(^{-1}\) (Girovich, 1996) occurs, which can lead to a temperature increase up to the boiling point. This temperature increase can be dangerous and can also lead to chemical urea hydrolysis. Therefore, we do not recommend using CaO for the prevention of urea hydrolysis in urine.

### 4.3. Advantages for inclusion in a nutrient recovery process

The stabilization of fresh urine with Ca(OH)\(_2\) has a number of advantages, especially when included in a process for nutrient recovery. Three of these advantages are described in the following sections.

#### 4.3.1. Ammonia stabilization for subsequent water removal

The stabilization of fresh urine strongly limits the amount of ammonia that will volatilize thus allowing for subsequent water removal techniques to be used. The reduction of the stabilized urine volume will reduce transport costs while the increasing salinity of the concentrated liquid will further reduce urease activity (Hotta and Funamizu, 2008).

One method to reduce the volume of urine is by vapor pressure distillation which Udert and Wächter (2012) showed was effective for nitrified urine. In the case of urine stabilized by Ca(OH)\(_2\), this method would not be applicable due to elevated temperatures during distillation and the subsequent loss of urea by chemical urea hydrolysis. Three possible alternative processes for volume reduction are electrodialysis, reverse osmosis or freeze/thaw concentration. However, besides urea stability, the energy demand will be decisive for the choice of the process (Udert and Wächter, 2012).
Another alternative process is passive evaporation at ambient temperatures (Bethune et al., 2015) which could be an effective method because the ambient temperatures would be ideal for preventing chemical urea hydrolysis, provided a saturated Ca(OH)$_2$ solution is maintained to prevent enzymatic urea hydrolysis. However, any post stabilization process, such as evaporation, should keep the temperature and pH within the recommended range during storage and transport. Consequently, any impacts on pH and temperature must be considered (in addition to energy requirements) when selecting downstream processing and distribution of stabilised urine. Moreover, other factors that should be considered include the ratio of monovalent to divalent cations in the stabilised urine which will affect soil sodicity and salinisation potential. Therefore, downstream processing would need to be evaluated on a case-by-case basis.

4.3.2. Potential targeted phosphate removal
The addition of Ca(OH)$_2$ does not only enable the preservation of urea, but also the separate recovery of phosphate by precipitation as calcium phosphate. Dosing a calcium salt for phosphate recovery by calcium phosphate precipitation is only economical if fresh urine with a high pH is used. If calcium is added to stored urine, that means after urea was hydrolyzed, calcium carbonate will precipitate instead of calcium phosphate.

To recover phosphate from stored urine, struvite can be precipitated by dosing a magnesium compound (Etter et al., 2011). However, a magnesium salt such as Mg(OH)$_2$ would not stabilize fresh urine even though it would result in struvite precipitation because simulation results (not shown) indicate the saturation pH would not exceed 10.

In addition, the use of Ca(OH)$_2$ over Mg salts has several advantages. The use of calcium phosphate in the phosphate industry is more likely compared to struvite (Driver et al., 1999). Calcium phosphate is less sensitive to heat, whereas inadequate drying, storing or transporting of struvite at temperatures higher than 60 °C can cause its decomposition and the loss of ammonia (Sarkar, 1991). Compared to Mg salts, Ca(OH)$_2$ is a widespread and low-cost material: costing approximately US$ 0.08 kg$^{-1}$ while magnesium chloride costs about five times as much as Ca(OH)$_2$ (Muster et al., 2012). Depending on where the Ca(OH)$_2$ is sourced from and the quality of the reagent, the typical cost of the chemical per person per day would equate to approximately US$ 0.0008 (assuming a dosage of 10 g Ca(OH)$_2$ L$^{-1}$ of fresh urine and a person producing 1 L of fresh urine per day).

4.3.3. Possible inactivation of pathogens
Literature shows that urine stabilization with the addition of Ca(OH)$_2$ and the corresponding pH increase could lead to pathogen inactivation. Alkaline treatment is a well-known practice in wastewater treatment for pathogen inactivation (EPA, 1978, 2000; Girovich, 1996). Farrell et al. (1974) showed that the presence of bacteria in wastewater sludge can be reduced to a negligible level at a pH higher than 11.5, with a recommended exposure of 0.5 h and more. The observed pathogen removal was even higher at pH values of 12.5, leading to a reduction of the required exposure time. At pH values of 11.5, complete inactivation of viruses in soil at an exposure time of less than 1 h was shown by Pancorbo et al. (1988). However, for the inactivation of parasitic nematode worms such as Ascaris, pH values higher than 12 and exposure to this highly alkaline environment for several months are required (Eriksen et al., 1996). While these literature reference suggest that the pH values in stabilized urine will at least kill viruses and pathogens, further investigations are needed to determine the potential of Ca(OH)$_2$ for disinfection.

5. Conclusions
We propose dosing Ca(OH)$_2$ as a new method for stabilizing fresh urine because (a) no active dosing is needed, (b) the resulting saturation pH is in an operating region where neither enzymatic or chemical urea hydrolysis is relevant and (c) bacteria and viruses are killed at these high pH values as previous literature suggests. We have shown that dosing with Ca(OH)$_2$ can also be used to precipitate calcium phosphate out of fresh urine as indicated by the decrease in dissolved phosphorus after the stabilization process. We also recommend providing a dosage of 10 g Ca(OH)$_2$ L$^{-1}$ of fresh urine. With this high dosage, saturation will be ensured for a wide range of urine compositions and operating conditions. Furthermore, we tentatively recommend an operating temperature range of 14 °C–40 °C to limit chemical urea hydrolysis.

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References


