Discussion

Corrigendum to "Solubility and speciation of iron in cementitious systems"

Fabio E. Furcas [a,*], Barbara Lothenbach [b], O. Burkan Isgor [c], Shishir Mundra [a], Zhidong Zhang [a], Ueli M. Angst [a]

[a] Institute for Building Materials, ETH Zürich, Stefano-Franscini-Platz 3, 8093, Zürich, Switzerland
[b] Empa, Concrete & Asphalt Laboratory, Switzerland
[c] School of Civil and Construction Engineering, Oregon State University, Corvallis, OR, USA

The authors regret to report a typo in the above article in Table 1. The heat of formation of $Fe(OH)_2^+$, as reported in row 13 should read $\Delta_f H_m^\circ = -550.2 \pm 1.0$ kJ mol^{-1}. Likewise, the heat of formation of FeO^+ in row 14 should read $\Delta_f H_m^\circ = -264.3 \pm 1.0$ kJ mol^{-1}.

The authors would like to apologise for any inconvenience caused and would like to highlight that the typos do not affect calculations presented throughout the paper, because the correct values were used in the calculations.

* Corresponding author.
  E-mail address: ffurcas@ethz.ch (F.E. Furcas).

Table 1

Summary of dependent reaction components as defined by the corrosion related hydrolysis and dissolution reactions enlisted in Table 2.

Component	MW, g mol-1	$\Delta_f G_m^{\circ}$, kJ mol^{-1}	$\Delta_f H_m^{\circ}$, kJ mol^{-1}	S_m°, J mol^{-1} K^{-1}	$C_{p,\,m}^{\circ}$, J mol^{-1} K^{-1}	V°, J bar^{-1}	Source
$Fe(OH)_2^+$	89.8597	-457.84 ± 0.87[b]	-550.2 ± 1.0	-11.8 ± 4.9			[1]: $\log_{10}\beta^{\circ}$, [2]: $\Delta_r H_m^{\circ}$
FeO^+ $(+H_2O(l) = Fe(OH)_2^+)$	71.8444	-220.74 ± 0.87[c]	-264.3 ± 1.0[c]	-81.8 ± 4.9	-200.9[a]	-4.20[a]	

[a] Retrieved via the HKF EOS [3].

[b] Internally computed from $\Delta_r G_m^{\circ} = \sum_i \nu_i \Delta_f G_{m,\,i}^{\circ}$.

[c] Obtained by subtracting $\Delta_f G_m^{\circ}(H_2O(l)) = -237.140 \pm 0.041$ kJ mol^{-1} and $\Delta_f H_m^{\circ}(H_2O(l)) = -285.830 \pm 0.040$ kJ mol^{-1} [4] from the previously enlisted species.

Declaration of competing interest

The authors declare the following financial interests/personal relationships which may be considered as potential competing interests: Fabio Enrico Furcas reports financial support was provided by ETH Zurich.

Data availability

Data will be made available on request.

References

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